**Freezing Point Depression and Boiling Point Elevation**

* When \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ dissolve in water, they break apart into their \_\_\_\_\_\_\_\_\_\_\_\_. These ions interfere with \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
* Adding salt \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ point and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ point.
* Different salts will break up into different \_\_\_\_\_\_\_\_\_\_\_\_\_ of ions.
* The \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ ions, the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the freezing point and the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the boiling point.

Q: Which salt, when mixed with H2O, will have the lowest freezing point?

1. NaCl b. MgCl2 c. FeBr3 d. LiBr

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